

Electrolysis

A Decomposition reaction.

Basic Setup

The primary concept for this lab is the interaction of Electrolysis of water from (Molecular form) H_2O to its (Elemental components) gases of Hydrogen H_2 and oxygen O_2 .

H_2 comes off at the cathode (negative) pole and O_2 comes of at the anode (positive).

Materials:

Graphite from Pencils

Beakers

Rubber bands

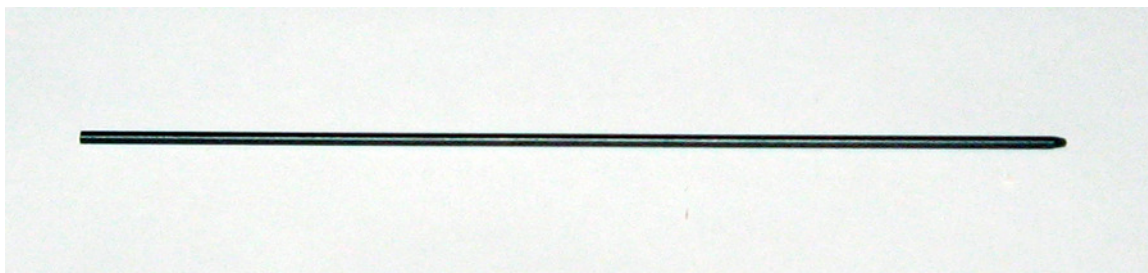
Knife (Maybe just the instructor)

Wire (solid core # 22)

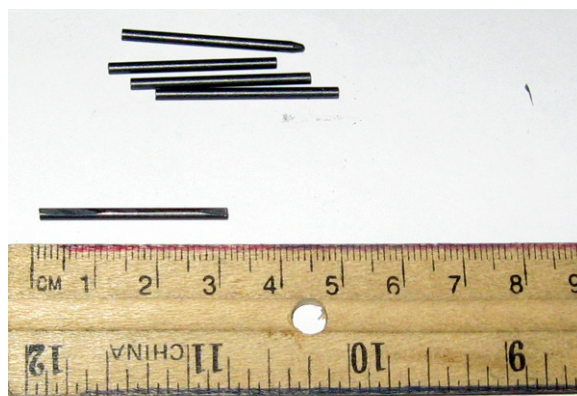
Hot Glue



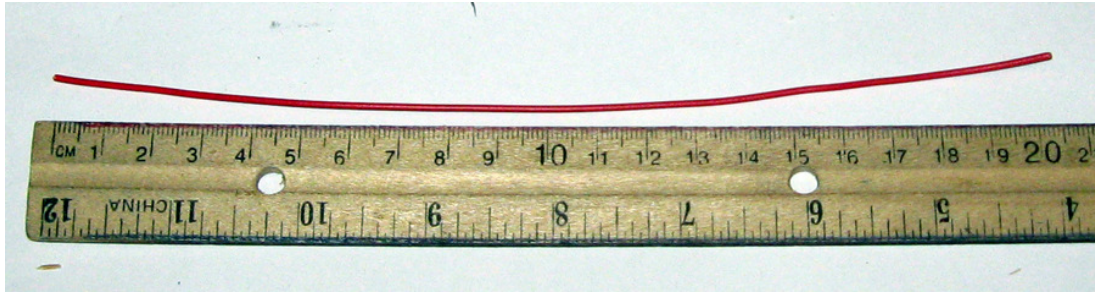
Whittle a pencil to expose the graphite core. (Your instructor may provide this to you)



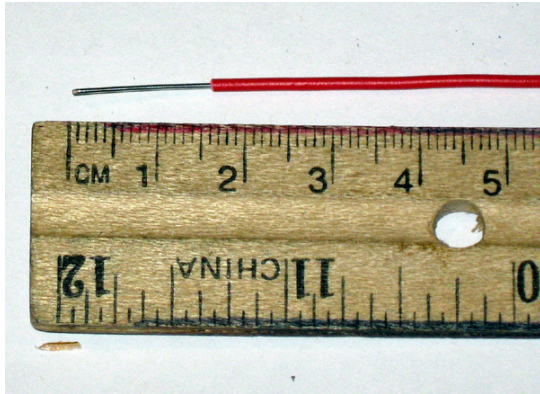
Carefully remove the graphite.



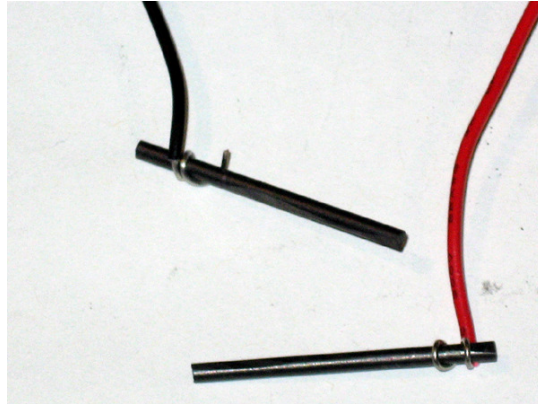
Break the graphite into pieces about 3 cm long.



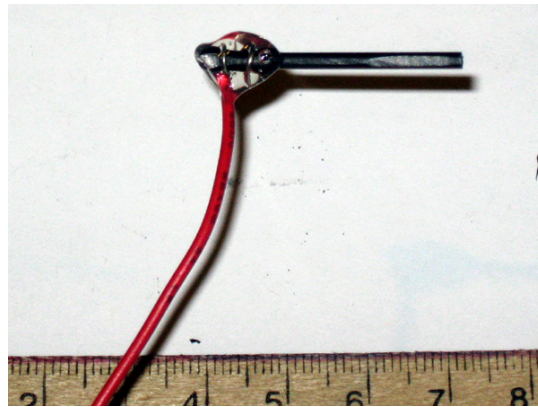
Cut a piece of black wire and red wire about 20cm long.



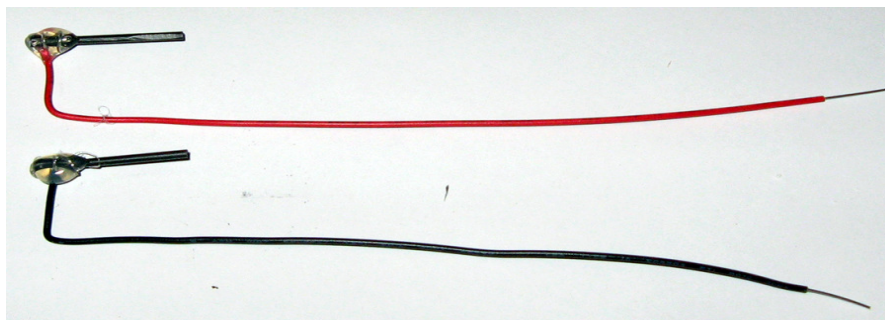
Strip about 1.5 cm off both wires



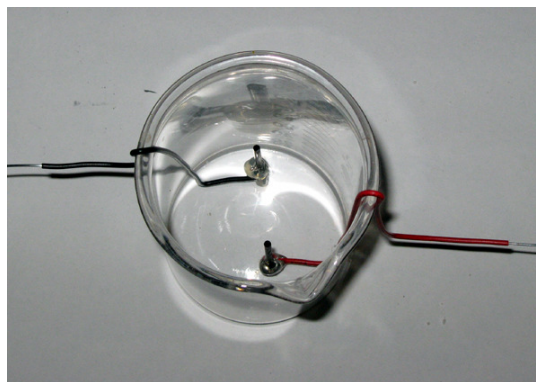
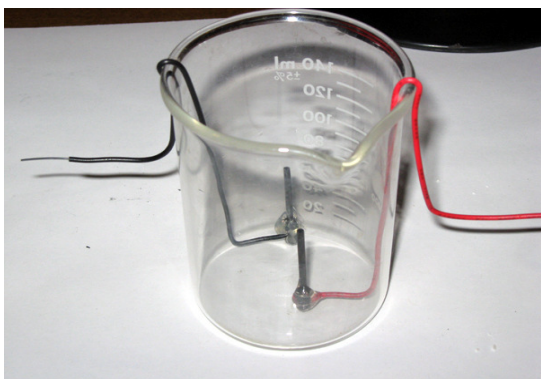
Carefully wrap the wires around the graphite pieces.



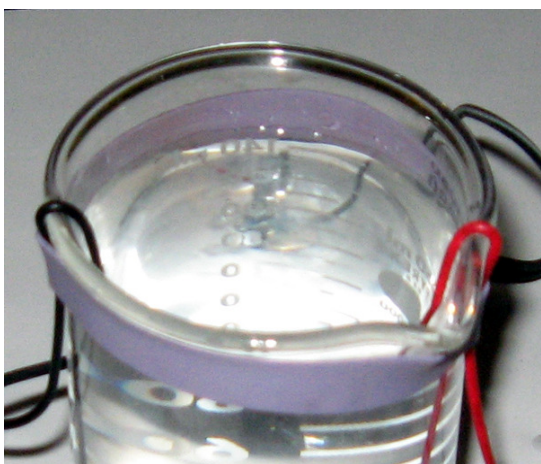
Place a glob of hot glue around the wrapped graphite rods.



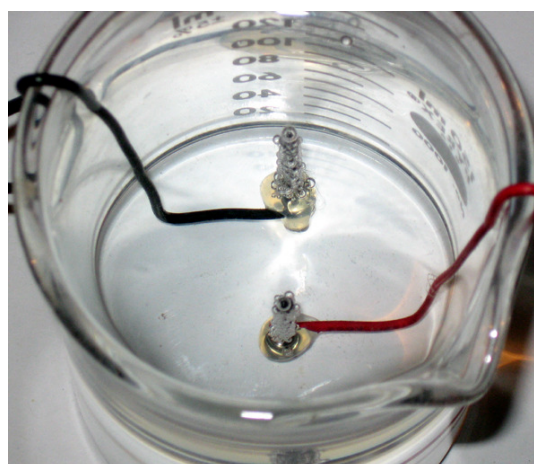
Bend the wires in a 90deg angle as shown.



Place the wires into a beaker and bend into place as shown.



Place a rubber band around the beaker to hold the wires firmly.



If all goes well you should have electrolysis
When you apply power!